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State The atomic

number is the same as

the number of protons

for each element From

the periodic table the

atomic number is the

whole number that is

used to put the

elements in order For

sodium it is 11 1

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For elements that have large atomic numbers, the most stable nuclei are those where $N/Z = 1.5$. The reason for the larger N/Z number is that neutrons are needed to stabilize the nuclei of heavier atoms. Some Rules to Help You Predict

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Nuclear Stability. 1. Except for and , all stable nuclei have a number of neutrons that is equal to or greater than the number of protons. 2. A nucleus that has an ...

Section 1 Atomic Nuclei and Nuclear Stability

SECTION 1 Name Class
Date Development of
the Atomic Theory
continued

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ph07ci_at000107a

Bohr Model of the Atom

In the Bohr model, electrons move around the nucleus like planets around the sun. What Is the Modern Atomic Theory? Atomic theory has changed over the past 100 years.

Scientists such as Erwin Schrödinger from Austria and

CHAPTER **Introduction to** **Atoms SECTION 1**

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Section 1 -

RADIOACTIVITY.

STUDY. PLAY. Identify
the two particles that
are located in the

nucleus. protons and
neutrons. The net
charge of the atomic
nucleus is __. positive .

Identify the three
forces that exist in the
nucleus. strong nuclear
force; electrical force;
weak force. Which
nucleon provides the

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nuclear strong force?
the neutron. The
nuclear strong force is
a(n) (attractive ...

Section 1 - RADIOACTIVITY Flashcards | Quizlet

What is a subatomic
particle with an atomic
mass of one and a 1
electric charge present
in all atomic nuclei?

Unanswered Questions
What is the basic
difference between DIV
and IDIV instruction in

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What is a subatomic particle with an atomic ... - answers.com

When atomic nuclei decay? We need you to answer this question! If you know the answer to this question, please register to join our limited beta program and start the conversation right now!

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**When atomic nuclei
decay - Answers**

11. Natural potassium consists of 93.1 % potassium-39 (atomic mass 38.964 u) and 6.9 % potassium-41 (atomic mass 40.962 u). Estimate and then calculate the average atomic mass of natural potassium. the average atomic mass of potassium should be close to the mass number of the most common isotope, so it

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should be about 39
amu

Answers to Review Questions for Atomic Theory Quiz #1

Section 4.1 studying
atoms 18 Terms.

maddy_mo_mo. PS 2

Match 1 24 Terms.

aliciayoung01. Atoms,
Elements, And Periodic
Table 62 Terms.

TeddyAdventurer.

OTHER SETS BY THIS

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Study Guide 75 Terms.

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Character Classes 22

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Chemistry Test 4 37

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Sports Medicine pt. 2

12 Terms.

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FOLDERS WITH ...

Section 4.1 Studying Atoms Flashcards | Quizlet

A positron is the
opposite of a β particle,
it is

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e^+ , the mass number will not change but the atomic number will decrease by 1. Answer e.

Electron capture has the same effect on the nucleus as positron emission: The atomic number is decreased by one and the mass number does not change.

3.1: Nuclear Chemistry and Radioactive Decay

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Quiz 3. STUDY.

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PLAY. Match . Gravity.

Created by. jesseking.

Conceptual Physical

Science Chapters

12-17. Terms in this set

(105) Atomic nucleus.

The dense, positively

charged center of

every atom. Electron.

An extremely small,

negatively charged

subatomic particle

outside the atomic

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nucleus. Element. Any material that is made up of only one type of atom ...

Quiz 3 Flashcards | Quizlet

Questions pertaining to atomic nucleus.

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questions (practice)
| Nuclei | Khan
Academy**

They determine the

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atomic number (Z) The atomic number determines the atom's identity Electronic cloud Nucleus A simplified view of the hydrogen atom, which consists of only one electron outside the nucleus. The nucleus contains only one proton and no neutrons. All other elements contain neutrons in their nuclei. 1. ATOMIC STRUCTURE

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FUNDAMENTALS

LEARNING OBJECTIVES

To review the basics ...

1. ATOMIC STRUCTURE FUNDAMENTALS

Section 1.3: Atomic Theory
What is a Theory? A theory is an explanation of observations that are supported by consistent, repeated experimental results. Some theories stand the test of time and

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become accepted.

Others are discarded or modified as new or conflicting evidence from experimental data or observations are presented. Ex. Theory of Evolution, Theory of Plate Tectonics What is a Law? A ...

Section 1.3: Atomic Theory - menihek.ca

14.1 Use a periodic table of the elements to answer these questions, 1. The

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following graphics represent the nuclei of atoms. Using a periodic table of elements, fill in the table. 2. How many protons and neutrons are in the nucleus of each isotope? a. hydrogen-2 (atomic number = 1) b. scandium-45 (atomic number = 21) c. aluminum-27 (atomic ...

14 1 Atoms And Isotopes Answer Key

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Answers...

History of the Atom -
Summary Sheet -
ANSWER KEY Scientist
Summary about Theory
Reasons for it's failure
continuous spectrum
Chadwick • showed
that the atomic nuclei
must contain HEAVEY
neutral particles as
well as positive
particles. • called
neutral subatomic
particles neutrons. N/A

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**History Of Atomic
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Answer Key

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25 Section Review
Problems Section
Review 25.1 Part A
Completion 1.
radioactive 2.
radioisotopes 3.

**ch.25 section review
answers - Answers**

Where To Download Atomic Nuclei Section 1. to Ch 25 Section ...

Section 25.3 Fission and Fusion of Atomic Nuclei 811 with ChemASAP Fission can be controlled so energy is released more slowly. Nuclear reactors, such as the one illustrated in Figure 25.11, use controlled fission to produce useful energy. In the controlled fission reaction within a nuclear reactor, much of the energy

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